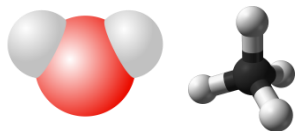
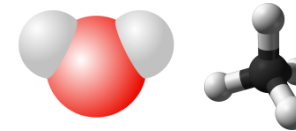


Name _____ Period _____



Covalent Bonding - Examples



| Atoms in the compound | Lewis Dot Diagram of the Final Compound | "Stick" Diagram |
|---|---|---|
| H_2 2 Hydrogen Atoms | $\text{H} \cdot \quad \text{H} \cdot \quad \rightarrow \quad \text{H} : \text{H}$ | $\text{H} - \text{H}$ |
| Cl_2 2 Chlorine Atoms | $:\ddot{\text{Cl}} \cdot \quad \cdot \ddot{\text{Cl}}: \quad \rightarrow \quad :\ddot{\text{Cl}} : \ddot{\text{Cl}}:$ | $:\ddot{\text{Cl}} - \ddot{\text{Cl}}:$ |
| HCl 1 Hydrogen + 1 Chlorine | $\text{H} \cdot \quad \cdot \ddot{\text{Cl}}: \quad \rightarrow \quad \text{H} : \ddot{\text{Cl}}:$ | $\text{H} - \ddot{\text{Cl}}:$ |
| NH_3 1 Nitrogen + 3 Hydrogen | $\begin{array}{c} \cdot \ddot{\text{N}} \cdot \\ \cdot \\ \text{H} \cdot \quad \text{H} \cdot \quad \text{H} \cdot \end{array} \quad \rightarrow \quad \begin{array}{c} \text{H} : \ddot{\text{N}} : \text{H} \\ \ddot{\text{H}} \end{array}$ | $\begin{array}{c} \text{H} - \ddot{\text{N}} - \text{H} \\ \\ \text{H} \end{array}$ |

Covalent Bonding – Questions

| | | |
|---|--|--|
| <p>CH₄</p> <p>1 Carbon + 4 Hydrogen</p> | | |
| <p>NO₃</p> <p>1 Nitrogen + 3 Oxygen</p> | | |
| <p>HF</p> <p>1 Hydrogen + 1 Fluorine</p> | | |

| | | |
|---|--|--|
| <p>H₂O</p> <p>2 Hydrogen + 1 Oxygen</p> | | |
| <p>CF₄</p> <p>1 Carbon + 4 Fluorine</p> | | |
| <p>Br₂</p> <p>2 Bromine Atoms</p> | | |
| <p>SF₂</p> <p>1 Sulfur + 2 Fluoride</p> | | |



2 Hydrogen +
1 Sulfur

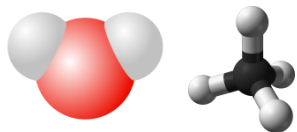


1 Nitrogen +
3 Bromine

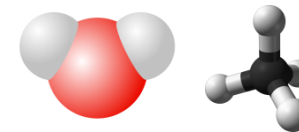


1 Carbon +
4 Bromine

Name _____ Period _____



Covalent Bonding - Examples



| Atoms in the compound | Lewis Dot Diagram of the Final Compound | "Stick" Diagram |
|-------------------------------------|---|-----------------|
| H_2 2 Hydrogen Atoms | | |
| Cl_2 2 Chlorine Atoms | | |
| HCl 1 Hydrogen + 1 Chlorine | | |
| NH_3 1 Nitrogen+ 3 Hydrogen | | |